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Chapter 16 – Acid-Base Equilibria. 16.1 Acids & Bases: A Brief Review. - Arrhenius acids and bases: -- acid: an  $H^+$  donor  $HA \rightleftharpoons H^+(aq) + A^-(aq)$  -- base: an  $OH^-$  donor  $MOH \rightleftharpoons M^+(aq) + OH^-(aq)$  - Brønsted-Lowry acids and bases: -- acid: an

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Michael Aronov

$H^+$  donor  $HA \rightleftharpoons H^+ + A^-(aq)$   
 $(aq) \rightleftharpoons (aq) + (aq)$  -- base: an  
 $H^+$  acceptor  $HB \rightleftharpoons B^-(aq)$   
 $(aq) \rightleftharpoons (aq) + (aq)$  .

## **Chapter 16 Acid- Base Equilibria - Directory**

BASES. Table 16.1. 1:  
General Properties of  
Acids and Bases.

produce a piercing pain  
in a wound. give a  
slippery feel. taste  
sour. taste bitter. are  
colorless when placed  
in phenolphthalein (an

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indicator). are pink when placed in phenolphthalein (an indicator). are red on blue litmus paper (a pH indicator).

## **16.1: Acids and Bases - A Brief Review - Chemistry LibreTexts**

16.10: Acid-Base Behavior and Chemical Structure Inductive effects and charge delocalization significantly influence

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the acidity or basicity of a compound. The acid-base strength of a molecule depends strongly on its structure. The weaker the A-H or B-H<sup>+</sup> bond, the more likely it is to dissociate to form an  $\text{H}^+$  ion.

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WORKSHEET~ 1. What

two substances are  
always produced by a  
neutralization reaction?

a. acid and a base b.

water and a base c.

water and an acid d.

water and a salt 2.



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## **Chapter 16: Leader: Acid and Base Review**

Chapter 16. 16.1 Acids and Bases: A Brief Review; 16.2 Bronsted-Lowry Acids and Bases; 16.3 The Autoionization of Water; 16.4 The pH Scale; 16.5 Strong Acids and Bases; 16.6 Weak Acids; 16.7 Weak Bases; 16.8 Relationship Between  $K_a$  and  $K_b$ ; 16.9 Acid-Base Properties of Salt

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16.11 Lewis Acids and  
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Chapter 16 Acids and  
Bases 1. Acids were  
recognized primarily  
from their sour taste.  
Bases were recognized  
from their bitter taste  
and slippery feel on  
skin. 2. In the  
Arrhenius definition, an  
acid is a substance  
that produces

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hydrogen ions ( $H^+$ )  
when dissolved in  
water, whereas a base  
is a substance that  
produces hydroxide  
ions ( $OH^-$ ) in

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## Section 16.1 – ACIDS AND BASES: A BRIEF REVIEW • Acids and

bases were first recognized by the properties of their aqueous solutions. For example, acids turn litmus red, whereas bases turn litmus blue.

- Arrhenius recognized that properties of acidic solutions are due to  $H^+(aq)$  ions and those of basic solutions are due to  $OH^-(aq)$  ions.

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