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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent: $[\text{Au}(\text{CN})_2]^-$, LaCl_3 , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by

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comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

Chapter 12.2: Stoichiometry of Reactions in Solution ...

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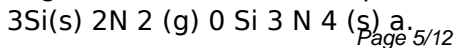
calculate the number of moles of each reactant present.

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1 Data Analysis
Data Analysis 1. A sample of aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample.
2. Saturn is about 1 429 000 km from the Sun.

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Chemistry: Supplemental Stoichiometry Problems. Directions:
Solve each of the following problems. Assume excess of any
reactant that isn't mentioned, unless otherwise . specified. Show
your work, including proper units, to earn full credit. 1. Given the
balanced equation, show what the following molar ratios (i.e.,
the conversion factors ...

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stoichiometry problem, the first step is always to calculate the

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number of moles Page 6/33 Chapter 12 Stoichiometry Practice Problems Chapter 12: Stoichiometry study guide by leahrosner includes 30 questions covering vocabulary, terms and more.

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